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2. (a) The rate
of the reaction

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will increase
because the rate
depends on the
concentration of

A as given in
the rate law:

$$\text{Rate} = k [A]^2.$$

the rate

constant is

independent of

the

concentration of

the reactants

and will not

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change. (b) The rate of the reaction will not change.

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kinetics can be divided into two parts. The first, at the macroscopic level, is the study of rates of reactions: what the rate of reaction means; how to determine a rate by

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experiment; and
how factors,
such as the
concentrations
of reactants and
temperature,
influence rates.

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KINETICS
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Determining rate
Answers
law from Initial
Rates. (Use the
ratio of initial
rates to get the
orders). 2.

Consider the
table of initial
rates for the
reaction: $2\text{ClO}_2 + 2\text{OH}^- \rightarrow \text{ClO}_3^- + \text{ClO}_2^- + \text{H}_2\text{O}$

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2 0. Experiment
[ClO₂]₀, mol/L
[OH⁻]₀, mol/L
Initial Rate,
mol/(L · s) 1
0.050 0.100 5.75
x 10⁻²

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constants, they
find rate-
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steps of
reactions, they
find half-life's
of reactions and
they determine
the overall
order of
reactions.

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based on the
concepts in AP
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Enduring
understanding
4.A: "Reaction
rates that

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depend on
temperature and
other

environmental

factors are
determined by
measuring

changes in
concentrations
of reactants or
products over
time." Group: AP

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5. If a reaction
proceeding by
the mechanism $A +$
 $B \rightarrow C +$ Doocurs at

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a rate x , and if
the

concentrations
of A and B are

both doubled,
what will be the
new rate of
reaction ...

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Kinetics: the branch of chemistry that deals with rates of reactions- chemistry & engineering.

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Worksheet 1-3 -
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Mechanisms Page
2 d) Given that
the uncatalyzed
overall reaction
is a slow
exothermic
reaction, draw a
potential energy

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graph which shows the possible shape of the curve for the uncatalyzed reaction. On the same graph, show a possible curve for the catalyzed reaction.

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Worksheet 1-3 -

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